

BR: January 15th

Use your States of Matter Activity to answer the following questions.

- How would you define endothermic?
... exothermic? *L requires*
↳ gives off heat energy
- What are intermolecular forces?
attraction
- What characteristics do gases have?
lots of energy
no set shape
moves rapidly
no color
fills container

Jan 22-12:42 PM

Tuesday January 15th

Objective: Determine the conditions of KMT (kinetic molecular theory).

- Bellringer
- States of Matter & Phase Changes ?'s
- Gas Properties Notes
- Phase Changes and KMT Online Activity

DUE: States of Matter activity

HW: Phase changes and KMT online Activity

Jan 21-8:46 AM

Go over States of Matter and Phase Changes			
	Solid	Liquid	Gas
Kinetic Energy of Particles	<i>low</i>	<i>med.</i>	<i>high</i>
Intermolecular Forces	<i>high</i>	<i>med</i>	<i>low</i>

Phase Change	Name	Inter forces	kinetic energy	exo/ endo
s → l	<i>melting</i>	<i>D</i>	<i>I</i>	<i>endo</i>
l → g	<i>evaporation</i>	<i>D</i>	<i>I</i>	<i>endo</i>
g → s	<i>Sub.</i>	<i>I</i>	<i>D</i>	<i>exo</i>
s → g	<i>Sub</i>	<i>D</i>	<i>I</i>	<i>endo</i>
g → l	<i>condensation</i>	<i>I</i>	<i>D</i>	<i>exo</i>
l → s	<i>freezing</i>	<i>I</i>	<i>D</i>	<i>exo</i>

Jan 22-4:00 PM

KMT: Kinetic Molecular Theory

- Gas particles are in constant, random motion.
- Gas particles are separated by relatively large distances.
- When gas particles collide, they do not transfer kinetic energy.
- Gas particles have no attractive or repulsive forces between them.
- The kinetic energy of a gas is dependent on the temperature of the gas.

Jan 24-1:07 PM

Assumptions:

- Gas particles are in constant, random motion
- Collisions between gas particles are completely elastic (no energy is lost when particles collide)
- Kinetic energy is the same for all gases at the same temperature

Jan 14-4:30 PM

PROPERTIES OF GASES

- Low density: few particles per volume
- Compression: squeeze particles into smaller space
- Expansion: move to fill space
- Diffusion: from high concentration to low concentration
- Effusion: moving from a puncture to a vacuum

Jan 26-7:50 AM

Phase Changes and Kinetic Molecular Theory

Go to preparatorychemistry.com/KMT_flash.htm

Complete questions 2-6 with your table partner.

Make sure you answer them in COMPLETE SENTENCES.

We will go over the first one as a class.

Jan 23-7:51 AM