

BR: April 30th

1. Kinetic energy is directly proportional to... *temp.*
2. What is the volume at STP given .875 moles? *PV=nRT*
3. How can you decrease entropy? *disorder ↓ #, solid*
4. Which substances dissolve in water? *ionic, polar, covalent*
5. Adding solute does what to boiling point and what to freezing point? *↑*
6. How can you increase the solubility of a gas? *↓T*
7. Find the percent by mass given the solute is 40 g and solution is 125 g

$$\frac{\text{solute}}{\text{solution}} \times 100 = \frac{40\text{g}}{125\text{g}} \times 100$$

Tuesday, April 30th

Objective: Students will review reaction types and naming.

1. Bellringer
2. Notes: Review Naming and Reaction Types
3. Worktime

HW: Naming and Writing Product Practice
DUE: Missing Work

May 2-8:40 AM

Naming:

- cation first, anion second
- cation name stays the same
- anion name ending changes to --ide, UNLESS it's a polyatomic ion
- polyatomic anions stay the same
 - > polyatomic anions: negative ion with more than 1 element

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Naming:

Name the following:

1. $(\text{NH}_4)\text{Cl}$ ammonium chloride
2. $\text{Fe}(\text{NO}_3)_3$ iron(III)nitrate
3. KI potassium iodide
4. Mg_3P_2 magnesium phosphide
5. $\text{Ca}(\text{OH})_2$ calcium hydroxide

Apr 28-8:11 AM

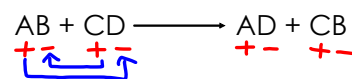
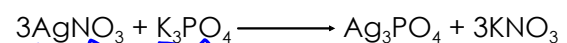
Writing Formulas:

Write the formula for the following:

1. sodium chloride NaCl
2. magnesium hydroxide $\text{Mg}(\text{OH})_2$
3. hydrogen sulfide H_2S
4. barium sulfate BaSO_4
5. ammonium hydroxide $\text{NH}_4(\text{OH})$

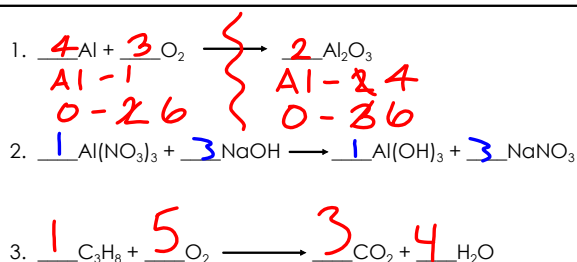
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Double Replacement Reactions:

General Form:Example:

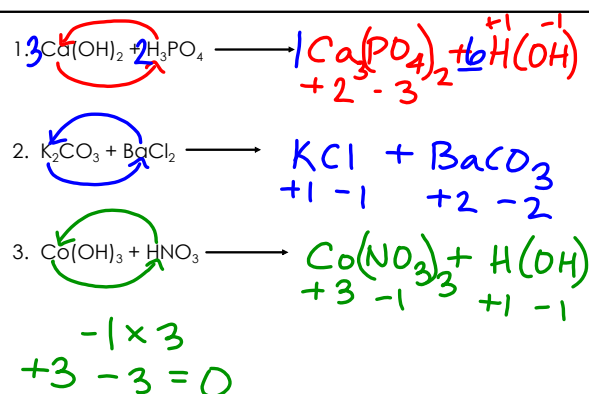
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Balancing:



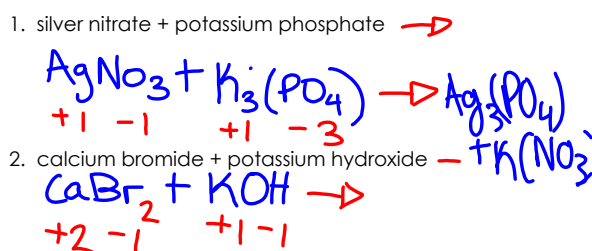
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Predicting Products:



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Measuring pH

- pH paper:** ^{indicators} will change the color depending on the hydrogen ion concentration in solution, the color is then compared to a standard scale
- pH meter:** more accurate than pH paper, contains electrode that are immersed in solution, will give a digital readout

May 5-8:09 AM