

Formula Writing Binary Ionic Compounds

-- You don't need to draw the transfer of electrons to write binary ionic compounds.

** all compounds are neutral!!!*
 What do you notice about the charge and the subscripts when you write the formula with Al and O?

e⁻ transfer

ions

Oct 5-10:44 AM

Formula Writing Binary Ionic Compounds

1. Figure out the charges for each ion
2. Drop the sign of the charge (+ or -)
3. Switch the number of the charge from the cation to the anion and vice versa.

Mg and P

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Practice:

1. Sr and Cl
 $SrCl_2$ ← $Sr^{+2} Cl^{-1}$
2. K and N
 K_3N ← $K^{+1} N^{-3}$
3. Be and N
 Be_3N_2 ← $Be^{+2} N^{-3}$
4. Mg and O
 MgO ← $Mg^{+2} O^{-2}$

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Naming Binary Ionic Compounds:

Cation: lose electrons, positive charge, metal
 -- NAME DOES NOT CHANGE

Anion: gain electrons, negative charge, nonmetal
 -- CHANGE THE ENDING TO --ide

example: K and S potassium sulfide

Oct 5-10:48 AM

Practice:

1. CaCl_2
Calcium Chloride
2. Rb_3N
Rubidium Nitride
3. Sr_3N_2 strontium Nitride
4. MgS magnesium Sulfide

Oct 5-10:47 AM