

## Electron Configurations:

· Sometimes called **ground state** electron configurations because all electrons are in their **lowest** possible energy levels

· Steps for writing a ground state electron configuration:

- Start from **hydrogen**
- Write **energy level** (1-7) (for s and p sublevel it is the period #)
- Write **sublevel** (s,p,d,f)
- Write number of **electrons** in the sublevel as an **exponent** (superscript)
- Stop at the desired number of **electrons**

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## Practice

Write the electron configuration for the following elements:

- Cobalt (Co)

- ~~Tungsten (W)~~ <sup>#47 Silver</sup> Ag

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## Practice

· Elements can be identified by their ending electron configurations

· Examples:

3d<sup>8</sup> Ni

2p<sup>4</sup> O

5s<sup>1</sup> Rb

1s<sup>2</sup> He

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